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Key indicators

Single-crystal X-ray study T = 97 K Mean σ (C–C) = 0.003 Å R factor = 0.044 wR factor = 0.122 Data-to-parameter ratio = 17.1

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e. The crystal structure of the title compound, $C_{29}H_{39}N_3O_4S_4,$ was determined at 97 K and possesses normal geometric parameters.

2-Methyl-N,N'-bis[2-(methylsulfanyl)ethyl]-

2-(2-pyridyl)-N,N'-ditosylpropane-1,3-diamine

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Comment

Organic ligands containing pyridyl and thioether functional groups have been used to study the copper coordination chemistry relevant to bio-inorganic systems (Klein *et al.*, 2004; Nonoyama *et al.*, 1975; Nonoyama, 1975; Karlin *et al.*, 1979; Champloy *et al.*, 1998). Recent efforts in our laboratory have focused on generating a family of ligands that contain a pyridyl group with two thioether arms (Klein *et al.*, 2005). We report here the crystal structure of the tosylated thioether amine ligand 2-methyl-N,N'-bis[2-(methylsulfanyl)ethyl]-2-(2-pyridyl)-N,N'-ditosylpropane-1,3-diamine, (I).



The title compound was synthesized by the deprotonation of N-(2-methylsulfanylethyl)-p-toluenesulfonamide with sodium hydride and subsequent addition of CH₃C(2-C₅H₄N)(CH₂OTs)₂ to form the tosylated amine product.

The molecular structure of (I) possesses normal geometric parameters. Intermolecular interactions between molecules are limited to polar and van der Waals interactions.

Experimental

N-(2-methylsulfanylethyl)-*p*-toluenesulfonamide (Mizukami & Kono, 1970) (6.618 g, 27.01 mmol) was dissolved in dry DMF under N₂ and heated to 378 K. Sodium hydride (2.72 g, 113.4 mmol) was added to this solution with stirring at 378 K. Stirring at 378 K was continued for 1 h following the cessation of hydrogen gas evolution. The reaction mixture was cooled to room temperature and the DMF solution was filtered from excess sodium hydride. $CH_3C(2-C_5H_4N)(CH_2OTs)_2$ (Friedrich *et al.*, 1997) (6.416 g, 13.50 mmol), dissolved in DMF, was added to this solution over a period of 2 h and stirred at 378 K for an additional 2 h. The reaction mixture was cooled to room temperature and poured into water. The aqueous solution was extracted with chloroform. The organic extracts were

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View of the title complex, showing the atom-labeling scheme. H atoms have been omitted for clarity. Displacement ellipsoids are drawn at the 50% probability level.

combined and the chloroform removed in vacuo to reveal a redbrown oil. The title compound was obtained from this crude product as a white powder following recrystallization from ethanol (2.46 g, 29.3% yield). Crystals suitable for X-ray structural analysis were obtained by slow evaporation of ethanol solutions of the title compound. (m.p. 406–409 K). ¹H NMR (300 MHz, CDCl₃): δ 8.45– 8.49 (m, 1H), 7.59–7.66 (m, 5H), 7.36–7.41 (m, 1H), 7.22–7.26 (m, 4H), 7.10–7.15 (m, 1H), 3.70 (d, J = 14 Hz, 2H), 3.39 (d, J = 14 Hz, 2H), 2.43–2.70 (*m*, 8H), 2.36 (*s*, 6H), 1.74 (*s*, 6H), 1.55 (*s*, 3H). ¹³C{H} NMR (75 MHz, CDCl₃): δ 163.13, 149.34, 143.82, 137.25, 136.29, 130.04, 127.68, 122.51, 122.35, 57.82, 50.51, 46.90, 31.75, 21.77, 19.80, 15.52. Analysis calculated for C₂₉H₃₉N₃O₄S₄: C 56.01, H 6.32, N 6.76%; found: C 55.99, H 6.30, N 6.59%.

Crystal data

$C_{29}H_{39}N_3O_4S_4$	Z = 2
$M_r = 621.87$	$D_x = 1.306 \text{ Mg m}^{-3}$
Triclinic, $P\overline{1}$	Mo $K\alpha$ radiation
a = 11.446 (2) Å	Cell parameters from 8357
b = 12.166 (2) Å	reflections
c = 12.584 (2) Å	$\theta = 2.3 - 28.2^{\circ}$
$\alpha = 92.881 \ (3)^{\circ}$	$\mu = 0.34 \text{ mm}^{-1}$
$\beta = 112.681 \ (3)^{\circ}$	T = 97 (2) K
$\gamma = 99.595 \ (3)^{\circ}$	Block, colorless
V = 1581.5 (5) Å ³	$0.47 \times 0.33 \times 0.31 \text{ mm}$
Data collection	
Bruker APEX CCD area-detector	6174 independent reflections

diffractometer ω scans Absorption correction: multi-scan (SADABS; Sheldrick, 2002) $T_{\min} = 0.857, T_{\max} = 0.902$ 17215 measured reflections

Refinement

5852 reflections with $I > 2\sigma(I)$ $R_{\rm int} = 0.023$

 $\theta_{\rm max} = 26.0^{\circ}$ $h = -14 \rightarrow 14$ $k = -14 \rightarrow 14$ $l = -15 \rightarrow 15$

$$\begin{split} &w = 1/[\sigma^2(F_{\rm o}^2) + (0.07P)^2 \\ &+ 1.42P] \\ &where \ P = (F_{\rm o}^2 + 2F_{\rm c}^2)/3 \\ (\Delta/\sigma)_{\rm max} = 0.001 \\ \Delta\rho_{\rm max} = 0.87 \ {\rm e} \ {\rm \AA}^{-3} \\ \Delta\rho_{\rm min} = -0.49 \ {\rm e} \ {\rm \AA}^{-3} \end{split}$$



Figure 2

The molecular packing of the title complex, viewed along the *a* axis. H atoms have been omitted for clarity. Displacement ellipsoids are drawn at the 50% probability level.

Table 1 Selected geometric parameters (Å, °).

S3-C18	1.802 (2)	N2-C8	1.470 (2)
S3-C17	1.808 (2)	N2-C16	1.479 (3)
S4-C29	1.799 (3)	N3-C19	1.472 (2)
S4-C28	1.806 (2)	N3-C27	1.492 (2)
C18-S3-C17	100.42 (11)	C7-C6-C19	109.87 (15)
C29-S4-C28	98.76 (11)	C19-C6-C8	101.91 (14)
C8-N2-C16	118.57 (16)	N2-C8-C6	115.19 (15)
C8-N2-S1	116.49 (13)	N2-C16-C17	114.24 (16)
C16-N2-S1	118.99 (13)	C16-C17-S3	111.55 (14)
C19-N3-C27	116.72 (15)	N3-C19-C6	115.02 (15)
C19-N3-S2	116.48 (12)	N3-C27-C28	111.47 (15)
C27-N3-S2	115.53 (12)	C27-C28-S4	111.05 (13)
C9-S1-N2-C8	-64.34 (16)	C27-N3-C19-C6	-101.63 (18)
C9-S1-N2-C16	88.16 (16)	C1-C6-C19-N3	57.0 (2)
C20-S2-N3-C19	64.97 (14)	C7-C6-C19-N3	-67.7(2)
C20-S2-N3-C27	-77.65 (14)	C8-C6-C19-N3	174.47 (15)
C8-N2-C16-C17	61.6 (2)	C19-N3-C27-C28	119.73 (17)
N2-C16-C17-S3	176.98 (14)	N3-C27-C28-S4	177.34 (13)
C18-S3-C17-C16	-75.92 (18)	C29-S4-C28-C27	-86.41 (16)
010 00 011 010	(10)2	010 010 010	00111 (1

H atoms were positioned geometrically (C-H = 0.95-0.99 Å) and refined using a riding model, with $U_{iso} = 1.2$ (1.5 for methyl) times $U_{eq}(C)$.

Data collection: SMART (Bruker, 1998); cell refinement: SAINT (Bruker, 1998); data reduction: SAINT; program(s) used to solve structure: SHELXTL (Sheldrick, 2000); program(s) used to refine structure: SHELXTL; molecular graphics: SHELXTL; software used to prepare material for publication: SHELXTL.

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References

- Bruker (1998). SMART and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
- Champloy, F., Benali-Cherif, N., Bruno, P., Blain, I., Pierrot, M., Reglier, M. & Michalowicz, A. (1998). *Inorg. Chem.* 37, 3910–3918.
- Friedrich, S., Schubart, M., Gade, L. H., Scowen, I. J., Edwards, A. J. & McPartlin, M. (1997). *Chem. Ber. Recl*, **130**, 1751–1759.
- Karlin, K. D., Dahlstrom, P. L., Stanford, M. L. & Zubieta, J. (1979). J. Chem. Soc. Chem. Commun. pp. 465–467.
- Klein, E. L., Anderson, B. M., Michels, J. T., O'Malley, B. G., Pal Chaudhuri, U., Grohmann, A. & Houser, R. P. (2005). Submitted.
- Klein, E. L., Khan, M. A. & Houser, R. P. (2004). Inorg. Chem. 43, 7272-7274.
- Mizukami, A. & Kono, M. (1970). JP Patent No. 4 5017 424.
- Nonoyama, M. (1975). Inorg. Chim. Acta, 13, 5-10.
- Nonoyama, M., Tomita, S. & Yamasaki, K. (1975). Inorg. Chim. Acta, 12, 33-37.
- Sheldrick, G. M. (2000). *SHELXTL*. Version 6.10. Bruker AXS Inc., Madison, Wisconsin, USA.
- Sheldrick, G. M. (2002). SADABS. University of Göttingen, Germany.